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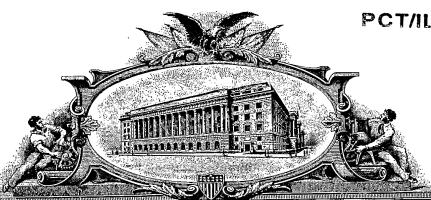
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May 27, 2004

2011 S Custos Crysta	South Cl mer Win al Plaza T	d Trademark Office ark Place dow, Mail Stop Patent Application Two, Lobby, Room 1B03 ginia 22202				
	Re:	New Patent Application in U.S. Applicant(s): Uri BANIN et al Title: NANOSUBSTRATE WITH CONDUCTIVE ZONE AND METHOD FOR ITS SELECTIVE PREPARTION				
	*	Atty's Docket: BANIN4A				
Sir: Atjāci	hed here	with is the above-identified application for Letters Patent including:				
1	[X]	Specification (17 pages), claims (4 pages) and abstract (1 page)				
	[X]					
	[X]	7 Sheets Drawings (Figures 1A-8) [X] FORMAL [] Informal				
	[X]	The inventors of this application are: Uri BANIN Taleb MOKARI				
	[] [X]	Information Disclosure Statement with SB/08A and references Return Receipt Postcard (in duplicate)				
The f	ollowing	statements are applicable:				
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	[X]	The present application claims the benefit of U.S. Provisional Appln. No. 60/541,248, filed February 4, 2004, and U.S. Provisional Appln. No. 60/554,913, filed March 22, 2004.				
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[X]	Amend the specification by inserting before the first line the sentence: []This is a continuation / division / continuation-in-part of copending parent application No							
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NANOSUBSTRATE WITH CONDUCTIVE ZONE AND METHOD FOR ITS SELECTIVE PREPARTION

FIELD OF THE INVENTION

This invention relates to the field of treatment of semiconductor nanostructures.

5 LIST OF REFERENCES

The following references are considered to be pertinent for the purpose of understanding the background of the present invention:

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The above references will be acknowledged in the text below by indicating their numbers [in brackets] from the above list.

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BACKGROUND OF THE INVENTION

Anisotropic growth of nanomaterials has led to the development of complex and diverse nano-structures such as rods, tetrapods, prisms, cubes and additional shapes. These architectures display new properties and enrich the selection of nano-building blocks for electrical, optical and sensorial device construction. Even greater complexity and new function is introduced into the nanostructure by anisotropic growth with compositional variations. This has been elegantly realized by growing semiconductor heterostructures such as p-n junctions and material junctions in nanowires [1, 2], and in the case of colloidal nanocrystals, in growth of a dot-rod of two different semiconductors [3] and in complex branched growth. In these examples, anisotropic growth was performed with the same material type (semiconductor).

A process for the preparation of nanocrystalline semiconductors, having rod-like shape of controlled dimensions is described in US 5,505,928 [13] and in WO 03/097904 [4] for especially Group III-V semiconductors, [4]. Nanocrystal particles having core with first crystal structure, and at least one arm with second crystal structure are described in WO 03/054953 [5].

Recently there have been several reports relating to connectivity formation 30 for micron-long quasi-one-dimensional structures such as nanotubes and

nanowires [6, 7, 8]. However, wiring in solution, of shorter semiconductor nanoparticles such as rods and tetrapods, with arm lengths of less then 100 nm is a difficult open problem.

5 SUMMARY OF THE INVENTION

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There is a need in the art for new nanostructures having selective, well-defined anchor points (preferably conductive anchor points) grown upon them for use in self-assembly in solution and onto substrates. Such nanostructures and method for their manufacture are not available to date.

Examples of desired nanostructures would be metal dots grown onto the tips of nanoparticles, in a controllable and repeatable manner that would also provide an electrical contact point. The conductive zones grown onto the tips of nanoparticles would provide well-defined anchor points onto which selective chemistries could be used to generate self-assembled structures of controlled arrangements.

The present invention thus provides in a first aspect new nanoscale materials in which a metal tip (conductive zone) is present on the edges of a nanostructure. The novel materials of the invention are nanostructures having an elongated shape such as rod, bipod, tripod and tetrapod. Excluded from the scope of the present invention are nanotubes and nanowires bearing electrodes formed by evaporation, such as those described in references [6-8] above.

The nanostructures of the invention are composed of at least one elongated structure element and comprise a first material, where an elongated structure element of the nanostructures bears an electrically conductive zone made of a second material.

The first material mentioned above is selected from semiconductor material, insulating material, metallic material and mixtures thereof. More preferably, the first material is a semiconductor material selected from Group II-VI semiconductors, Group IV-VI semiconductors, Group IV-VI semiconductors, Group IV semiconductors, alloys made of these semiconductors, combinations of

the semiconductors in composite structures and core/shell structures of the above semiconductors. Even more preferably, the nanostructures are made from Group II-VI semiconductors, alloys made from Group II-VI semiconductors and core/shell structures made from Group II-VI semiconductors. Specific examples 5 of Group II-VI semiconductors are CdSe, CdS, CdTe, alloys thereof, combinations thereof and core/shell layered-structures thereof.

The second material mentioned above is a metal or metal alloy. Preferably, the metal is a transition metal. Specific examples of such transition metals are Cu, Ag, Au, Pt, Co, Pd, Ni, Ru, Rh, Mn, Cr, Fe and Ti. In a preferred 10 embodiment, the first material is different than the second material.

The present invention provides, in another of its aspects, a method for forming such an electrically conductive zone on a nanostructure having at least one elongated structure portion. The method of the present invention is carried out in liquid medium and it comprises: contacting a solution comprising nanostructures with a solution comprising a metal or metal alloy source, to obtain upon isolation, nanostructures bearing an electrically conductive zone on said at least one elongated portion thereof. The reaction is carried out at a temperature between about -40°C to about 350°C, preferably between about 10°C to about 80°C, more preferably between about 20°C to about 30°C and even more preferably at room temperature.

According to a preferred embodiment the reaction is carried out in the presence of at least one of the following agents in addition to said nanostructures and metal source: electron donor, surfactant and stabilizer.

The nanostructures used in the method of the invention have an elongated shape, for example of rods, wires, tubes, or in branched form. More preferably the nanostructures have an elongated shape such as for example nanorods and branched shape such tripods, tetrapods and the like. The term "nanorod" or "rod" as used herein is meant to describe a nanoparticle with extended growth along the first axis while maintaining the very small dimensions of the other two axes, 30 resulting in the growth of a rod-like shaped nanocrystal of very small diameter in

the range of about 1 nm to about 100 nm, where the dimensions along the first axis may range from about several nm to about 1 micrometer. The term "tetrapod" is meant to describe a shape having a core from which four arms are protruding at tetrahedral angles. In the case of nanorods, the resulting structures 5 after treating them with the metal or metal alloy source are shaped as "nanodumbbells".

The nanostructures have an elongated shape or even a branched shape and serve as a template at the nanometer level for the deposition of a conducting material, and as it will be described and exemplified herein below, the deposition 10 is accomplished in a controllable manner on the edges of the elongated portions of the nanostructures.

The nanostructures are made of a material comprising semiconductor material, insulating material, metallic material or mixtures thereof. Preferably, the nanostructures are made of semiconductor material selected from Group II-15 VI semiconductors, such as for example CdS, CdTe, ZnS, ZnSe, ZnO and alloys (e.g. CdZnSe); Group III-V semiconductors such as InAs, InP, GaAs, GaP, InN, GaN, InSb, GaSb and alloys (eg. InAsP); Group IV-VI semiconductors such as PbSe and PbS and alloys; and Group IV semiconductors such as Si and Ge and Additionally, combinations of the above in composite structures alloys. of sections with different semiconductor materials, for example consisting CdSe/CdS or any other combinations, as well as core/shell structures of different semiconductors such as for example CdSe/ZnS core/shell nanorods, are also within the scope of the present invention.

The nanostructures may also be made of an insulating material such as for example oxides and organic materials or, alternatively the nanostructures are made of metals. Examples of oxides are silicon oxide, titanium dioxide, zirconia. Metals include Au, Ag, Cu, Pt, Co, Ni, Mn and the like, and various combinations and alloys thereof. Organic materials suitable for use in the nanostructures are for example polymers.

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The metal or metal alloy source used in the method of the present invention preferably comprises a transition metal or mixture of such metals. A variety of metals may be used. This includes noble metals such as Cu, Ag, Au, or other transition metal elements such as Pt, Co, Pd, Ni, Ru, Rh, Mn, Cr, Fe, Ti and the like. The metal growth procedure is done by using a proper metal salt source, for example AuCl₃ for Au growth, Ag(CH₃COO) for silver growth, Cu(CH₃COO)₂ for Cu growth, PtCl₄ or Pt(acetylacetonate) for Pt growth, Ni(cyclooctadiene)₂ for Ni growth, Co₂(CO)₈ or CoCl₂ for Co growth, and Pd(NO₃)₂ for Pd growth.

The metal salts are dissolved in a proper organic solvent such as hydrocarbons, e.g. hexanes, cyclohexanes, etc., aromatic solvents e.g. toluene, etc., using a proper surfactant and/or stabilizer that stabilizes the nanostructures and the metal salt by preventing aggregation. The organic solvent used in the method of the present invention is one capable to solubilize both the nanostructure and the metal source.

Examples of surfactants are cationic surfactants such as ammonium salts, alkyl pyridinium and quaternary ammonium salts. More specific examples are tetrabutylammonium borohydride (TBAB), dodecyldimethylamonium bromide (DDAB), cetyltrimethylammonium bromide (CTAB), and salts of quaternary ammonium with acetate group ions such as acetate group ions, pivalate, glycolate, lactate and the like.

Stabilizer compounds used in the method of the invention are such compounds capable to coordinate to the nanostructure surface and/or the metal particle surface and hence prevent aggregation of the nanostructures. Examples of stabilizers are aliphatic amines, e.g. hexadecylamine, dodecylamine, octylamine, alkylthiols, e.g. hexane thiol, decylthiol, dodecylthiol, etc. and carboxylic stabilizers such as myristic acid, palmitic acid and citrate.

The metal or metal alloy salt is first dissolved in an organic solvent comprising a surfactant and a stabilizer to give a solution which is subsequently

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added in a controllable manner and a suitable temperature to the nanostructures solution.

When an electron source is desired in the method of the invention, an electron donor compound may be used. Examples of electron donors are organic 5 compounds, such as aliphatic amines, hydrides such as sodium borohydride and the like, ascorbic acid and other reducing agents. According to another example the electron source is obtained from an electron beam device. Alternatively, one may use electromagnetic radiation in order to excite the nanoparticles or the metal source.

More specifically, the present invention provides a method for forming in solution medium an electrically conductive zone on a nanostructure having at least one elongated structure, the method comprising: contacting an organic solution comprising semiconductor nanostructures with an organic solution comprising a metal or metal alloy source, a stabilizer and/or surfactant to obtain upon precipitation semiconductor nanostructures bearing at least one electrically conductive zone on said at least one elongated structure thereof. Preferably, the nanostructures used in the method of the invention are in the form of nanorods, tetrapods or any other branched structure and are made of elements of Group II-VI, alloys of such elements or core-shell layered structures thereof.

The method of the present invention provides new functionalities to the nanostructures, the most important of which is the formation of anchor points for directed self assembly. The selective tip growth of metal contacts provides the route to an effective wiring scheme for soluble and chemically processable nanostructures with branched shapes. This would allow to fully realize the 25 potential of miniaturization of devices using such nano-building blocks, while employing the powerful principles of self-assembly to connect them to the 'outside' world.

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BRIEF DESCRIPTION OF THE DRAWINGS

In order to understand the invention and to see how it may be carried out in practice, preferred embodiments will now be described, by way of non5 limiting example only, with reference to the accompanying drawings, in which:

- Fig. 1 illustrates TEM (Transmission Electron Microscopy) images A-D showing growth of Au onto CdSe quantum rods of dimensions 29x4 nm (length x diameter). Fig. 1A shows the rods before Au growth, Fig. 1B shows selective Au growth of about 2.2 nm; Fig. 1C shows selective Au growth of about 2.9 nm; and Fig. 1D shows selective Au growth of about 4 nm.
 - Fig. 2A illustrates the EDS (Energy dispersive X-ray spectroscopy) spectrum of a goldenized CdSe rod sample. The relative atom percentage of Au:Cd:Se is18%:42%:40%.
- Fig. 2B illustrates a powder X-ray diffraction comparing CdSe rods before 15 (1), and after (2) Au growth.
 - Figs. 2C and D illustrate HRTEM images of a single nano-dumbbell and a nano-dumbbell tip, respectively. The CdSe lattice for the rod in the center, and Au tips at the rod edges, can be identified as marked.
- Fig. 3 illustrates TEM images A-H showing growth of Au on tips of various CdSe quantum rods and CdSe tetrapods. Figs. 3A and 3B: 12x4 nm quantum rods before and after Au growth, respectively. Figs. 3C and 3D: 29x4 nm quantum rods before and after Au growth, respectively. Figs. 3E and 3F: 60x6 nm quantum rods before and after Au growth, respectively. Au growth on CdSe tetrapods showing a general view is presented in Fig. 3G and higher magnification image for one tetrapod is presented in Fig. 3H.
 - Fig. 4A illustrates the absorption spectra for CdSe/ZnS core/shell nanorod sample with varied Au tip size compared to the original rod template, where Au tip size is indicated for each trace. Spectra are offset vertically for clarity. Inset shows TEM image of the sample after Au growth (scale bar is 50 nm).
- Fig. 4B illustrates the photoluminescence (PL) spectra for CdSe/ZnS 01530047\11-01

core/shell nanorod sample with varied Au tip size compared to the original rod template, where Au tip size is indicated for each trace. Traces were multiplied by 25, 50 and 50 for the 2 nm, 3.2 nm, and 4.5 nm Au tips, respectively, for clarity. Inset shows a plot of relative PL yield for template (Φ₀) over Au-rod (Φ), versus Au ball size. Measurements were performed for rod solutions in a sealed cuvette under Ar using the 454 nm line of an Ar-ion laser with intensity of 100 mW. Fluorescence was collected using identical conditions for all solutions in a right angle configuration with a spectrograph/CCD setup, with 500 ms integration time.

- Fig. 5 shows sizing histograms for goldenized rods shown in Fig. 1. Histograms for rod diameter (Fig. 5A), length (Fig. 5B) and Au diameter (Fig. 5C) are shown for the four samples: 1. Original 29x4 nm rods, 2. Rods (10 mg) after treatment with 4 mg AuCl₃, 25 mg DDAB and 40 mg dodecylamine. 3. Rods (10 mg) after treatment with 8 mg AuCl₃, 50 mg DDAB and 90 mg dodecylamine. 4. Rods (10 mg) after treatment with 13.5 mg AuCl₃, 100 mg DDAB and 160 mg dodecylamine.
- Fig. 6 illustrates TEM of product from mixture of rods with AuCl₃ and DDAB without dodecylamine, Fig. 6A. before exposure to the TEM electron beam aggregated rods are seen. Fig. 6B. after exposure to the TEM electron beam Au patches appear on the rods.
 - Fig. 7 shows conductive atomic force microscopy current image of single nano-dumbbell measured at a sample bias of 2V. Higher conductance through the Au tips is observed, as seen also by the current cut taken along the rod (along plotted line). In the inset, a TEM image with the same scale of a Au-rod is shown for comparison.
 - Fig. 8 shows the self assembly of nano-dumbbells into chains formed by adding hexane dithiol bifunctional linker to a solution of nano-dumbbells.

DETAILED DESCRIPTION OF THE INVENTION

The method is exemplified hereinbelow with reference to selective growth of metal tips onto semiconductor nanorods and tetrapods.

In a method for selective growth of contacts made of gold, AuCl₃ was 5 dissolved in toluene by use of dodecyldimethylamonium bromide (DDAB) and dodecylamine, and the resulting solution was added to a toluene solution comprising of colloidal grown nanorods or tetrapods. The method is exemplified for the prototypical CdSe nanocrystal system that is highly developed synthetically and widely studied for its size and shape dependent properties.

CdSe rods and tetrapods of different dimensions (see below), were prepared by high temperature pyrolisys of suitable precursors, in a coordinating solvent containing a mixture of trioctylphosphineoxide (TOPO), and of In a typical Au growth reaction, a gold solution was phosphonic acids [9]. prepared containing 12 mg AuCl₃ (0.04mmol), 40 mg of DDAB (0.08mmol) and 70 mg (0.37mmol) of dodecylamine in 3 ml of toluene and sonnicated for 5 minutes at room temperature. The solution changed color from dark orange to light yellow. 20 mg of CdSe quantum rods of the required dimensions were dissolved in 4 ml toluene in a three neck flask under argon. The gold solution was added drop-wise over a period of three minutes. During the addition, carried 20 out at room temperature, the color gradually changed to dark brown. Following the reaction, the rods were precipitated by addition of methanol and separated by centrifugation. The purified product could then be redissolved in toluene for further studies.

Fig. 1 presents transmission electron microscopy (TEM) images showing growth of Au onto CdSe quantum rods of dimensions 29x4 nm (length x diameter). Fig. 1A shows the rods before Au growth, while in Figs. 1B-D, selective Au growth onto the rod tips is clearly identified as the appearance of points with enhanced contrast afforded by the higher electron density of the Au compared with CdSe. The rods now appear as 'nano-dumbbells'. Moreover, by controlling the amount of initial Au precursor, it is possible to control the size of

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the Au tips on the nano-dumbbell edges, from ~2.2 nm in Fig. 1B, to ~2.9 nm in Fig. 1C, to ~4.0 nm in Fig. 1D as summarized in Table 1. The procedure clearly leads to the growth of natural contact points on the tips of the rods.

5 Table 1

Sample	NC's ¹	HDA ²	DDAB ³	AuCl ₃	Rods size	Gold ball
	amount(mg)	amount(mg)	amount(mg)	amount(mg)	(Lx D)nm	size (nm)
1	-	-	-	-	29 x 4 nm	(original rod)
2	10 mg	40 mg	25 mg	4 mg	25.6 x 3.3 nm	2.22 nm
3	10 mg	90 mg	50 mg	8 mg	23.9 x 3.4 nm	2.9 nm
4	10 mg	160 mg	100 mg	13.5 mg	20.8 x 3.2 nm	4 nm

¹ NC - nanocrystals

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An additional observation from the analysis of ~200 particles per sample is that the overall rod length becomes shorter upon Au growth, and there is also a decrease in the average diameter of the rods, (Table 1 and Fig. 5 for the complete sizing histograms). Control experiments with the DDAB and dodecylamine without AuCl₃ were carried out and also in that case the average rod dimensions decreased, implying that reduction of rod size is perhaps related to dissolution of rods in the presence of DDAB and not to the Au growth.

Several structural and chemical characterization methods have been carried out in order to verify the material content and structure of the gold on the rod tips. Fig. 2A shows EDS analysis of a micron area of rods after growth and the appearance of Au in the goldenized and purified rod sample is clear. The powder Xray-diffraction pattern for the 29x4 nm rod sample comparing the rods

² HDA – hexadecylamine

³ DDAB – dodecyldimethylamonium bromide

before and after gold growth is shown in Fig. 2B. The appearance of the Au (111), (200) and (220) peaks following Au growth is evident, demonstrating crystalline Au is formed on the tips.

Further evidence for Au growth onto single rods, is provided by HRTEM (high resolution TEM) studies of the nano-dumbbells. Fig. 2C shows a HRTEM image of a single rod after gold treatment. The lattice image for the rod part composed of CdSe corresponds to growth of rods along the CdSe <001> axis. The Au is discerned once again as the region at the edge with enhanced contrast and the gold lattice is also shown in Fig. 2D.

Relating to the interface at the Au-CdSe, it is suggested that Au-Se bonds are formed, analogous to the known AuSe material [10]. This means that the interface is formed with covalent chemical bonds between the metal and the semiconductor and hence can be expected to provide good electrical connectivity.

The method for selective Au growth could be easily expanded and applied to rods of arbitrary dimensions, and to tetrapods, as well as to growth of other metals and to rods made of various semiconductor materials.

Fig. 3 shows TEM images for three rod samples of dimensions 12x4 nm (Fig. 3A, B), 29x4 nm (Fig. 3C, D), and 60x6 nm (Fig. 3E, F), before and after 20 Au treatment. The presence of the high-contrast tips on the treated rods, forming nano-dumbbells, is evident in all cases. Highly selective tip growth is discerned and demonstrated for three rod sizes and could easily be applied to arbitrary rod sizes. In addition the method was applied to a CdSe tetrapod sample, as can be seen in Fig. 2G showing several tetrapods, and in Fig. 2H showing an enlargement of one tetrapod, following the Au growth process. In this case, the growth occurs selectively on all the tips of the tetrapods leading to a tetrahedral arrangement for the Au tips, and once again providing the natural contact points for this unique structure, for further self-assembly and for electrical connections.

In another example, CdTe nanostructures served as the template for growing various metals on its edges. The synthesis of the CdTe in different

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shapes is known [14]. In a typical synthesis of CdTe rods, a mixture of 1 mmol of CdO dissolved in 1.125 gr oleic acid and 2.5 gr of 1-octadecene is heated in three neck flask to 300°C to obtain a clear colorless solution. In the glove box, a solution of Te (0.5 mmol of Te is dissolved in 1 ml of TOP) is prepared and 5 brought out in a vial sealed with septum to the injection. After the injection of the Te solution into the mixture in the three necked flask, the mixture is cooled to 260°C for growth. Modification of this procedure in terms of the temperature or precursor concentration results in size and shape changes. The oleic acid is used as a ligand and it dissolves the CdO in the octadecene.

Another specific semiconductor material that may be used is CdS, which is controllable in size and shape. The synthesis is based on the same principle which is injection precursor to hot solution, the Cd and S precursor in this case is Cd(S2CNEt)2 that could be synthesized according to known literature method [12]. In typical synthesis of CdS nanorods, a warm solution of Cd(S₂CNEt)₂ (50 mg dissolved in about 0.3 g of hexadecylamine (HDA) at about 70°C) is injected into hot solution of HDA and after 1 hr is cooled to 70°C and treated with ethanol and separated by centrifuging. Controlling the shape of the nanocrystals is done by changing the growth temperature of the synthesis from 300°C (rods) to 120°C (tetrapods).

Metal tips by the method described above have also been grown onto CdSe/ZnS core/shell nanorods (29x4 nm rods with 2 monolayer ZnS shell) with initial emission quantum yield of 2% [15]. Treatment of these rods with DDAB and dodecylamine without Au led to an increased quantum yield of 4%, likely because of the effect of the excess amine. Several Au sizes were grown from 25 about 2 nm to about 4.5 nm Au at the tips of the rods.

The metalized structures (in the case of Au growth the formed structures are termed herein "goldenized structures") exhibit new and different electronic, electrical and optical properties as compared to the original rods, due to the strong effect of the metal on the semiconductor properties. Absorption and 30 photoluminescence (PL) measurements were carried out to study the effect of Au

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growth on the rod optical properties as shown in Fig. 4. Absorption spectra (Fig. 4A) for the small Au tips on the rods still shows the excitonic structure but with increased absorbance in the visible and the appearance of a tail to the red of the particle gap. Upon increased Au size, the features of the absorption of the rods 5 are washed out and the tail to the red becomes more prominent. The spectra should contain in principle contributions from the semiconductor part and the plasmon resonance associated with the Au tips. However, attempts to add spectra of the rod template and Au nanocrystals did not reproduce the observed absorption and we suggest that the spectra are not a simple sum of both 10 components because of the modified electronic structure of the Au-rod nanodumbbell system. The strong mixing of the semiconductor and metal electronic states leads to modified density of states exhibiting broadened levels and a reduced band-gap.

The significant coupling of the Au is also observed for the PL (Fig. 4B) 15 that undergoes considerable quenching with increased Au ball size, by a factor of about 100 initially for the smaller Au balls (about 2 nm), and gradually down to a factor of about 500 for the large Au balls (about 4.5 nm). Quenching of the emission by the metal edges is expected via the new non-radiative pathways created by the proximity of metals, likely leading to electron transfer to the Au. Moreover, a systematic dependence of quenching on Au size is seen as shown in the Stern-Volmer type plot (inset of Fig. 4B). Both absorption and emission spectra exemplify the significant effect of the Au on the semiconductor rod properties in this new system, further proving the strong bonding of the Au to the CdSe rod.

The selective tip growth of Au onto the rods and tetrapods not only provides important developments for enabling electrical connectivity and new paths for self-assembly for such nanostructures. It is also an interesting and novel chemical reaction route with clear selectivity and anisotropic character. The reaction mechanism for the gold growth entails a reduction of Au. 30 Examining by TEM the Au solution with DDAB and dodecylamine, already

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reveals the formation of Au particles. When the reaction is carried out without dodecylamine, considerable aggregation of the CdSe rods was seen (Fig. 6A). Additionally, without the amine, growth of Au on rods was not apparent initially and only after the irradiation under the electron beam of the TEM we could observe some Au growth (Fig. 6B).

One of the benefits of the method of the present invention is its specificity leading to selective tip growth. This results from the preferential adsorption of the metal, e.g. Au complex formed in the Au solution by adding Au salt to DDAB and dodecylamine onto the nanostructures edges. The tips are more 10 reactive due to the increased surface energy and also possibly due to imperfect passivation of the ligands on these faces, which also leads to preferential growth along the <001> axis of CdSe rods. Once Au nucleates on the edge, it is preferential for additional Au to adhere and grow on that seed. This gains support from controlling the extent of Au growth on the rod tips by using 15 increased concentration of Au in the gold solution as was shown in Fig. 1. Moreover, early Au growth as shown in Fig. 1B reveals that in some rods preferential early growth occurs on one tip, in agreement with the surfactantcontrolled growth model suggested for CdSe rods [9].

It is important to note that in some cases Au growth was identified on branching and defect points, but at slower rate compared to the distinctive tip growth discussed above. This can be seen in Fig. 3E and 3G, where weak dark Au spots appear also in some positions other then the tips of the long rods and tetrapods. This growth can be controlled by the amounts of Au added to the rods. At such defect points, such as points where the diameter of the structure changes, 25 there is also increased reactivity due to the imperfect chemical bonding and increased surface energy. This leads to Au adhesion and growth in agreement with the mechanism for tip growth. It is emphasized that the tip growth occurs more readily and rapidly then growth on the defects and hence can be controlled to achieve contact points.

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The method may easily be expanded to additional semiconductor nanocrystal systems and to additional metals, to tailor the metal tip contacts as desired and the semiconductor element as well.

One application for the metal tips is in serving as electrical contact points. The role to be played by the Au tips as contact points for wiring the rods is exemplified by conductive atomic force microscopy (C-AFM) measurements carried out on goldenized 60x6 nm rods. Rods were deposited onto a conducting highly ordered pyrolitic graphite substrate, and embedded in a thin layer of poly methyl methacrylate (PMMA) to avoid dragging by the tip as reported earlier for 10 regular rods [16]. The current image of a single rod measured by this method reveals that already at a bias of 1.5-2 V, small tunneling current is flowing through the tips which are composed of Au, while the central part of the rod consisting of the semiconductor is non-conductive at these conditions (see Fig. 7). The small tunneling current is determined by the tunneling barriers at tipnanocrystal and nanocrystal-substrate junctions, dominated primarily by the PMMA. This measurement reveals the significantly higher conductance of the Au tips which would be critical for using them as electrical contact points.

Several strategies can be employed to realize such contacts. It is possible to form the metallized nanorods or other branched structures onto a substrate, 20 identify their position, and then write by electron-beam lithography electrodes to overlap with the Au tips. In a different approach, it is also possible to deposit the metalized rods onto pre-existing electrode structures, with or without Since the metal tipped electrostatic trapping by an applied electric field. nanostructures enable the connectivity to electrode structures, this clearly opens 25 the path for using them as transistors, in sensing applications, andin light emitting or light detecting devices.

The metal edges can also impart the rods with advantageous and novel They exhibit enhanced linear and non-linear optical optical properties. properties. The polarizibility of such a structure may obviously be significantly 30 increased compared with that of the regular rods. For example, enhancement in

second harmonic generation and also the observation of novel plasmon resonances related to highly controlled distances that could be tailored for the metal tips on rods.

Additionally, is possible to apply the powerful approach of self assembly 5 by using for example, biological templates e.g. DNA, for creating the connections to the metal tips of nanorods or of branched structures, or bifunctional ligands such as dithiols or diamines for binding preferentially to the Au tips. In such applications the metal tips serve as selective anchor points for ligands and chemistries preferential for the Au surface. Such self assembly could 10 for example be done in solution or onto surfaces. In solution, examples include formation of AAAA chains where A represents rods of one type. This is done by adding bifunctional ligands such as dithiols, for example hexane dithiol, to a solution with goldenized nanorods. The preferential binding of thiols to the Au tips leads to chain formation as can be seen in Fig. 8. Another example is the formation of ABAB chains where A represents one rod type and B another rod type. Here biochemical linkers such as avidin-biotin chemistry or DNA linking can be used to make selective ABAB chains. In another approach, combining tetrapods with rods on tip to tip basis may lead to formation of propeller structures.

The same chemistries can be used to self-assemble rods and tetrapods with Au tips onto patterned or non-patterned substrates. For example, a gold or silicon substrate is used together with a bifunctional ligand that binds with one function to the substrate and with the second function to the Au tip on the nanostructure.

Metal tipped structures also provide selective metal growth points for additional materials via a seeded growth solution-liquid-solid mechanism.

CLAIMS:

- 1. Nanostructure composed of at least one elongated structure element and comprising a first material, an elongated structure element of said nanostructure bearing an electrically conductive zone made of a second material.
- 5 2. The nanostructure of claim 1, wherein said first material is selected from semiconductor material, insulating material, metallic material and mixtures thereof.
 - 3. The nanostructure of claim 2 wherein said first material is a semiconductor material.
- 10 4. The nanostructure of claim 3 wherein said semiconductor material is selected from Group II-VI semiconductors, Group III-V semiconductors, Group IV-VI semiconductors, Group IV semiconductors, alloys made of these semiconductors, combinations of the semiconductors in composite structures and core/shell structures of the above semiconductors.
- The nanostructure of claim 4 wherein said nanostructures are made from Group II-VI semiconductors, alloys made from Group II-VI semiconductors and core/shell structures made from Group II-VI semiconductors.
 - 6. The nanostructure of claim 1 wherein said second material is selected from metal and metal alloy.
- 20 7. The nanostructure of claim 6 wherein said metal is a transition metal.
 - 8. The nanostructure of claim 7 wherein said transition metal is selected from Cu, Ag, Au, Pt, Co, Pd, Ni, Ru, Rh, Mn, Cr, Fe and Ti.
 - 9. The nanostructure of claim 1 having an elongated shape selected from rod, bipod, tripod and tetrapod.
- 25 10. A method for forming in a liquid medium, an electrically conductive zone on a nanostructure having at least one elongated structure element, the method comprising: contacting a solution comprising nanostructures with a solution comprising a metal or metal alloy source, to obtain upon isolation nanostructures bearing at least one electrically conductive zone on said at least one elongated structure thereof.

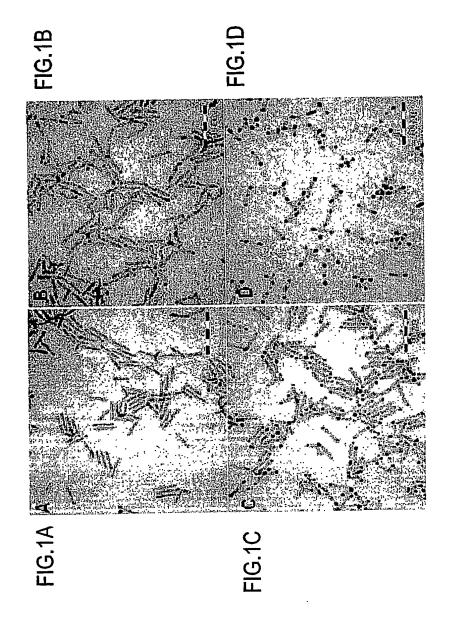
- 11. The method according to claim 10 wherein said nanostructure is made of a material comprising semiconductor material, insulating material, metallic materialor mixtures thereof.
- 12. The method according to claim 10 wherein said nanostructure is made of semiconductor material.
 - 13. The method according to claim 10 wherein said nanostructure has an elongated shape.
 - 14. The method according to claim 13 wherein said elongated shape comprises a branched shape.
- 10 15. The method according to claim 14 wherein said branched shape comprises rod, bipod, tripod and tetrapod.
 - 16. The method according to claim 12 wherein said nanostructure is made of a semiconductor material selected from Group II-VI semiconductors, Group III-V semiconductors, Group IV-VI semiconductors, Group IV semiconductors, alloys made of these semiconductors, combinations of the semiconductors in composite structures and core/shell structures of the above semiconductors.
 - 17. The method according to claim 16, wherein said nanostructures are made from Group II-VI semiconductors, alloys made from Group II-VI semiconductors and core/shell structures made from Group II-VI semiconductors.
 - 18. The method according to claim 11 wherein said nanostructure is made of an insulating material selected from oxides and organic polymers.
 - 19. The method according to claim 10 wherein the metal or metal alloy source solution further comprises a surfactant and/or a stabilizer.
- 25 **20.** The method according to claim 19 wherein said surfactant is a cationic surfactant.
 - 21. The method according to claim 19 wherein said stabilizer prevents aggregation of nanoparticles during the formation of an electrically conductive zone on a nanostructure.

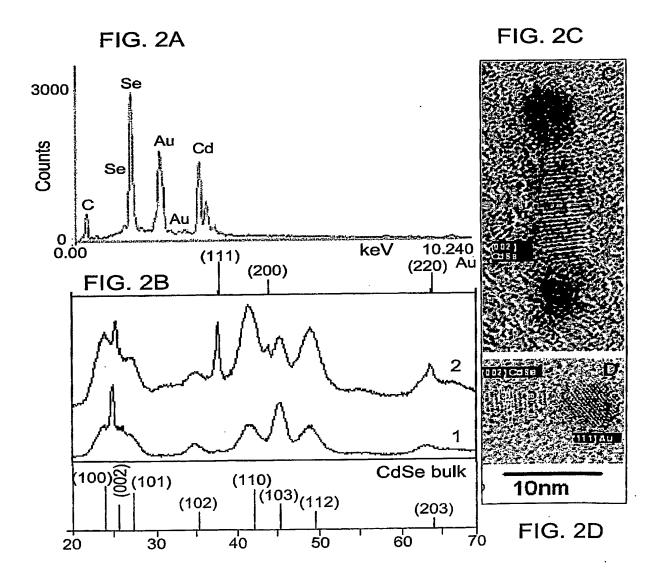
- 22. The method according to claim 20 wherein said stabilizer is selected from ammonium salts, alkyl pyridinium alts and quaternary ammonium salts.
- 23. The method according to claim 10 wherein said metal or metal alloy source comprises a transition metal element.
- 5 24. The method according to claim 23 wherein said metal or metal alloy source is a salt of a transition metal or transition metal alloy.
 - 25. The method according to claim 24 wherein said transition metal is selected from Cu, Ag, Au, Pt, Co, Pd, Ni, Ru, Rh, Mn, Cr, Fe and Ti.
- 26. The method according to claim 24 wherein said metal or metal alloy salt is first dissolved in an organic solvent comprising a surfactant and/or a stabilizer to give a mixture which is subsequently added in a controllable manner to the nanostructures solution.
 - 27. The method according to claim 10 wherein said electron donor is an organic compound.
- 15 28. The method according to claim 27 wherein said electron donor is selected from aliphatic amine, hydride and ascorbic acid.
 - 29. A method for forming in solution medium an electrically conductive zone on a nanostructure having at least one elongated structure element, the method comprising: contacting, an organic solution comprising semiconductor nanostructures with an organic solution comprising a metal or metal alloy source, a stabilizer and/or surfactant and/or electron donor to obtain upon precipitation
 - semiconductor nanostructures bearing at least one electrically conductive zone on said at least one elongated structure thereof.
 - 30. The method according to claim 29 wherein said nanostructures are in the form of nanorods, bipods, tripods or tetrapods.
 - 31. The method according to claim 29 wherein said semiconductor nanostructures are made of a material comprising elements of Group II-VI, alloys of such elements or core-shall layered structures thereof.

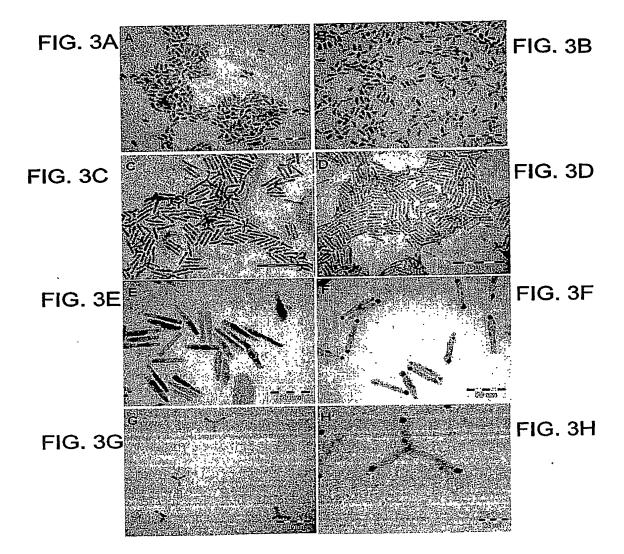
- 32. The method according to claim 31 wherein said semiconductor nanoparticles are made of a material comprising CdSe, CdS, CdTe, alloys thereof, combinations thereof or core/shell layered-structures thereof.
- 33. The method according to claim 29 wherein said electrically conductive zone comprises a metal selected from Au, Ag, Cu, Pt, Co, Pd, Ni, Ru, Rh, Mn, Cr, Fe, Ti or mixtures of such metals.
 - 34. Article of manufacture comprising the nanostructure of claim 1.
 - 35. An electronic device comprising the nanostructure of claim 1, or into which the nanostructure of claim 1 is integrated.
- 10 36. An electrode comprising the nanostructure of claim 1.
 - 37. An optical device comprising the nanostructure of claim 1, or into which the nanostructure of claim 1 is integrated.
- 38. Self assembled construct comprising a plurality of nanostructures according to claim 1, wherein each nanostructure is linked to another nanostructure in the construct through its conductive zone.

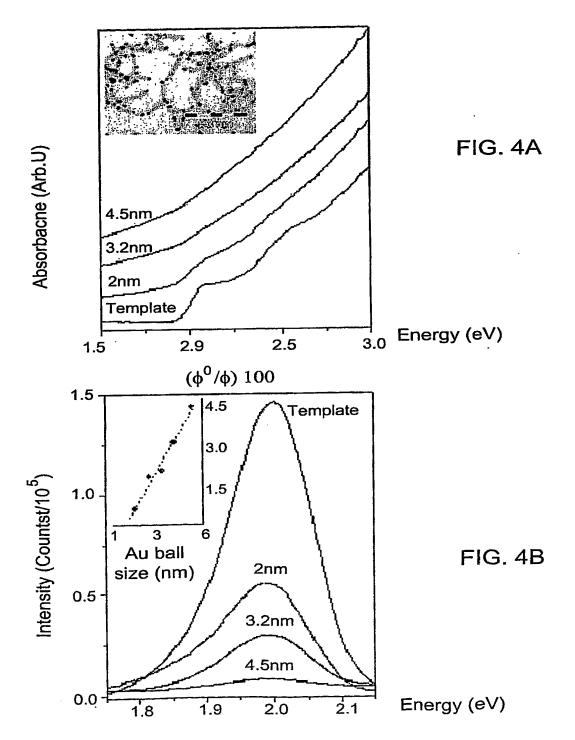
ABSTRACT

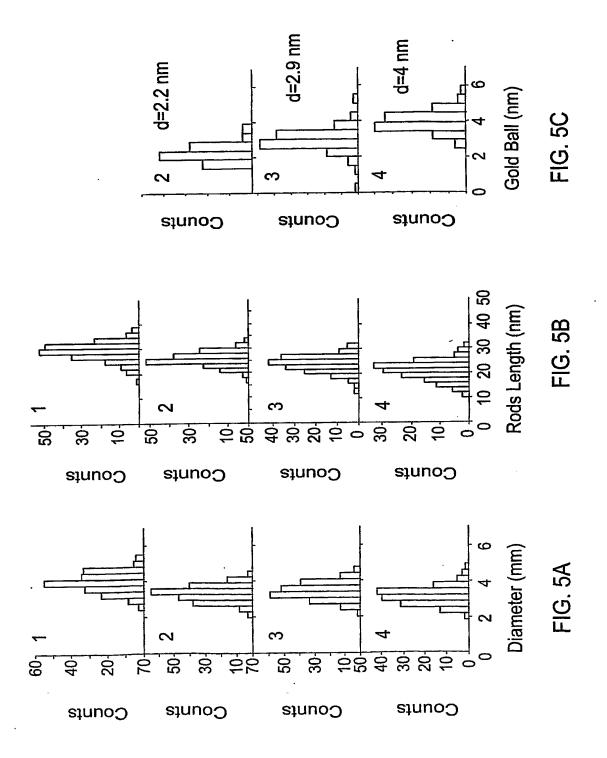
The present invention provides novel nanostructure composed of at least one elongated structure element, an elongated structure element of said nanostructure bearing an electrically conductive zone selectively grown onto the elongated structure element. The present invention further provides a selective method for forming in a liquid medium, such nanostructures.

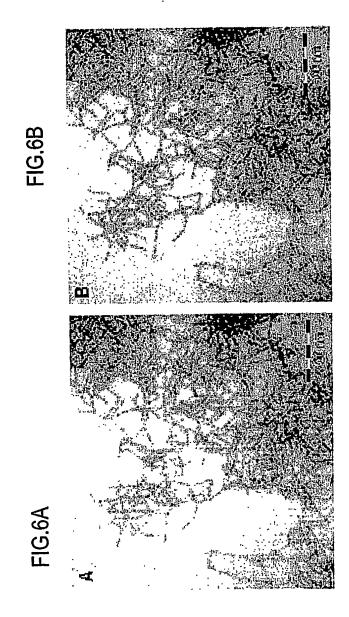












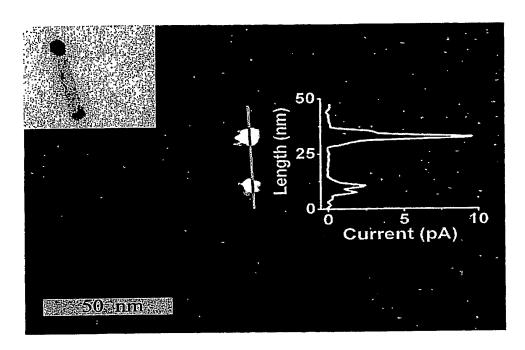


FIG.7

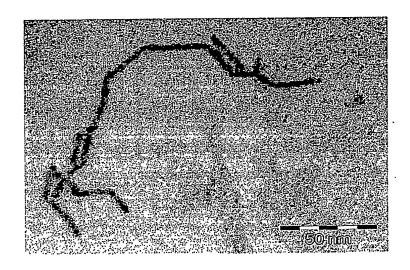


FIG.8